

C1 Topic 2 limestone revision

Quarrying	
1. State 1 advantage of quarrying	Get building materials, jobs
2. State 1 disadvantage of quarrying	Dust, noise, increase traffic, visual pollution, destroy habitats
3. How do you make concrete?	Mix cement, sand, water, gravel
4. How do you make mortar?	Mix cement, sand, water
5. How do you make cement?	Heat limestone with clay
Thermal decomposition (and reactions)	
6. What is limestone made out of?	Calcium carbonate
7. What type of reaction is this: calcium carbonate \rightarrow calcium oxide + carbon dioxide?	Thermal decomposition
8. What do you call a reaction where a compound breaks down due to being heated?	Thermal decomposition
9. Complete this equation: magnesium carbonate \rightarrow	Magnesium oxide + CO ₂
10. Complete this equation: zinc carbonate \rightarrow	Zinc oxide + CO ₂
11. Complete this equation: calcium oxide + water \rightarrow	Calcium hydroxide
12. What can calcium hydroxide be used for?	Neutralise acid/ limewater
13. What does limewater test for and what is the result?	CO ₂ , goes cloudy
Carbonates and acid rain	
14. Complete this equation: carbonate + acid \rightarrow	Salt + water + CO ₂
15. What can damage limestone?	Acid rain
16. Why are limestone statues and buildings damaged by rain?	Because it can be acid rain which damages limestone
Higher only	
17. What does calcium oxide react with to produce calcium hydroxide?	Water
18. What does calcium hydroxide react with to produce calcium carbonate?	Carbon dioxide